

Answer on Question #64050, Physics / Atomic and Nuclear Physics

How many neutrons does element X have if its atomic number is 58 and its mass number is 166?

Find: N - ?

Given:

Z=58

A=166

Solution:

The difference between the mass number (A) and the atomic number (Z) gives the number of neutrons (N) in a given nucleus: $N=A-Z$ (1)

Of (1) $\Rightarrow N=108$

Answer:

108

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