

Answer on Question Question #58126, Physics / Mechanics | Relativity

Task Calculate the electric power which must be supplied to the filament of a light bulb operating at 3000K. The total surface area of the filament is $8 \times 10^{-6} \text{ m}^2$ and its emissivity is 0.92.

Solution

According to black body theorem luminosity (or same, power), that going from surface, can be defined by formula:

$$L = \gamma A \sigma T^4$$

γ – emissivity, A – surface area, σ – constant, $5.67 \cdot 10^{-8}$, T – surface temperature

$$L = 0.92 \cdot 5.67 \cdot 10^{-8} \cdot 8 \cdot 10^{-6} \cdot 3000^4 = 41.7312 \cdot 10^{-14} \cdot 3^4 \cdot 10^{12} = 3380.2272 \cdot 10^{-2} \\ \approx 33.8 \text{ W}$$

Answer $L=33.8 \text{ W}$

Task Calculate the change in internal energy of 2kg of water at 90oC when it is changed to 3.30 m³ of steam at 100oC. The whole process occurs at atmospheric pressure. The latent heat of vaporization of water is $2.26 \times 10^6 \text{ J/kg}$.

Solution

According to thermodynamics $dU=PdV-TdS$, where dU is change in internal energy, P – pressure, dV – change in volume, TdS – heat, that given to system.

P is constant (according to condition) $dV=V_2-V_1$

$$V_1=2 \text{ liters}=0.002 \text{ m}^3 \quad V_2=3.3 \text{ m}^3$$

We can assume, that $dV=3.3 \text{ m}^3$

$TdS=Q=cm(t_2-t_1)+qm$, where $t_2=100 \text{ C}$ (before vaporization), $t_1=90 \text{ C}$ – initial temperature, q – latent heat of vaporization, c – heat capacity (4200 J/kg/C), m – mass of water.

$$Q = 4200 \cdot 2 \cdot 10 + 2 \cdot 2.26 \cdot 10^6 = 4.604 \cdot 10^6 = 4.6 \cdot 10^6$$

$$PdV = 100000 \cdot 3.3 = 330000$$

$$dU = PdV - Q = 0.33 \cdot 10^6 - 4.6 \cdot 10^6 = -4.27 \cdot 10^6$$

Answer

$$dU = -4.27 \cdot 10^6 \text{ J}$$