The question is: How many positive or negative charges are present in 300 g of water?

We have:

$$\vartheta(H_2O) = \frac{300g}{18\frac{g}{mol}} = 16.67mol$$

$$1mol = 6.02 * 10^{23} molecules$$

So,

$$300g = \frac{300}{18}mol = \frac{300}{18} * 6.02 * 10^{23} molecules = \frac{300}{18} * 6.02 * 10^{23} * 10 \ electrons$$

$$(because \ 1 \ molecule \ of \ water = 1 + 1 + 8 \ electrons \ or \ protons)$$

Result:

$$300g = 1.003 * 10^{26} \ electrons \ (or \ protons)$$