

According to the First law of thermodynamics:

$Q = c * m * \Delta T$, where: Q – heat energy, m – mass (liquid), c – heat capacity

$$\Delta T = \frac{Q}{c*m}$$

The greater rise in temperature had a liquid with a smaller heat capacity or (and) smaller mass.

The smaller heat capacity and smaller mass have ethanol:

C(ethanol)=2.46

C(water)=4,183

M(100ml ethanol)=0,08 Kg

M(100ml water)=0.1 Kg

Ethanol experiences the greater rise in temperature.