a)

It is impossible to produce water from calcium carbonate, but if you mean *what mass of water is necessary to treatment of 1 tonne of calcium carbonate to calcium hydroxide*, I found the solution (below).

The molar mass of calcium carbonate is $M(CaCO_3)=40+12+3*16=100$ g/mol

We have $n(CaCO_3) = m(CaCO_3)/M(CaCO_3) = 1000000g/100g/mol = 10000$ mol of calcium carbonate.

$CaCO_3(s)=CaO(s)+CO_2(g)$ (1)

$CaO(s)+H_2O(l)=Ca(OH)_2(s)$ (2)
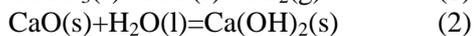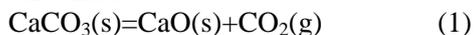
According to the equation (1) and (2) we need one mol of water for 1 mol of calcium carbonate, hence we have:

$n(H_2O) = n(CaCO_3) = 10000mol$

The mass of water is:

$m(H_2O) = n(H_2O)*M(H_2O) = 10000mol*18g/mol = 180000$ g $= 180$ kg

b) The mass of water that you need to add to the calcium oxide is equal to previous solution.

$M(water) = 180$ kg

c)

The molar mass of calcium carbonate is $M(CaO)=40+16=56$ g/mol

According to the equation (1) and (2) we need 1 mol of CaO for 1 mol of calcium carbonate, hence we have:

$n(CaO) = n(CaCO_3) = 10000mol$

The mass of CaO is:

$m(CaO) = n(CaO)*M(CaO) = 10000mol*56g/mol = 560000$ g $= 560$ kg