

- According to definition $\ln(K_2/K_1) = (E_a/R)\{(1/T_1) - (1/T_2)\}$
- Rate constant K_1 at temperature T_1
- Rate constant K_2 at temperature T_2
- E_a is the activation barrier. R is the Universal gas constant
- Here it is given that $K_1 = 0.0132 /s$, $K_2 = 0.684/s$, $T_1 = 400K$, $T_2 = 450K$
- Through Calculation The activation energy will be 118.158 KJ

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