

Question # 86134, answer

You need to make an aqueous solution of 0.221 M chromium(III) acetate for an experiment in lab, using a 500 mL volumetric flask. How much solid chromium(III) acetate should you add?

Answer:

- 1) Number of moles of chromium(III) acetate needed to make a solution $n = M \times V = 0.221$ moles/L $\times 0.5$ L = 0.1105 moles
- 2) Mass of solid chromium(III) acetate = $n \times MW = 0.1105$ moles $\times 229.128$ g/mole = 25.32 g

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