

**Question #85051, Chemistry / General Chemistry**

A 2.5M aqueous solution of  $\text{Fe}_2(\text{SO}_4)_3$  is \_M in  $\text{Fe}^{3+}$  and \_M in  $\text{SO}_4^{2-}$

**Answer**

A 2.5M aqueous solution of  $\text{Fe}_2(\text{SO}_4)_3$  is **5M** in  $\text{Fe}^{3+}$  and **7.5M** in  $\text{SO}_4^{2-}$

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