

How can i prepare 50% of hydrogen peroxide from 35% hydrogen peroxide?

Solution.

For the $V(\text{H}_2\text{O}) = 100 \text{ ml}$

$$1. \quad 0.35 = \frac{x}{100+x}$$

$$X = 53.84 \text{ g} - m(\text{H}_2\text{O}_2)$$

$$2. \quad 0.5 = \frac{x}{100+x}$$

$$X = 100 \text{ g} - m(\text{H}_2\text{O}_2)$$

$$100 - 53.84 = 46.16 \text{ g}$$

Answer. Add 46.16 g of H_2O_2 per 100 ml of water.

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