

how many moles are contained in 10.5 L of oxygen?

Solution. We have 10.5 L oxygen. At the same time, the conditions are not specified, therefore, we accept under normal conditions. Then the number of moles of oxygen will be: $n_{O_2} = \frac{V_{O_2}}{V_m} = \frac{10.5}{22.4} = 0.47$ moles.

Answer: 0.47 moles.

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