

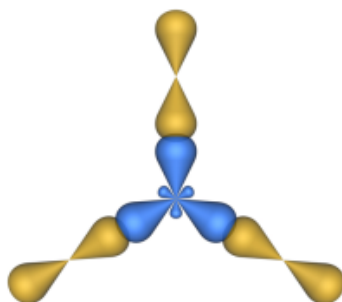
Question #83956, Chemistry / Inorganic Chemistry | for completion

Describe how the hybridization states of boron in the compound boron trichloride and show how the boron-chloride bond may be formed

Answer:

The spatial configuration of the molecule, the central atom of which includes sp^2 -hybrid orbitals.

This type of hybridization is observed, for example, in the BCl_3 molecule. The model of this molecule is shown in the figure:



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