

## Question #83805, Chemistry / General Chemistry | for completion

An important part of the Haber - Bosch Process is a Catalyst of iron with a small amount of aluminium added. explain how this catalyst changes the rate of the Haber-Bosch Process

Answer:

The Haber-Bosch process operates at high pressure so as to shift the equilibrium to the right, and high temperature to increase the rates of the reaction. Of course, operating at high temperature actually shifted the reaction to the left, but the trade-off for faster rates was accepted. By removing the ammonia as liquid ammonia, the equilibrium is continuously shifted to the right.

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