

Answer on Question #83575 – Chemistry – General Chemistry

Explain the difference in boiling point between F_2 ($-188^\circ C$) and HCl ($-85^\circ C$)

Solution:

F_2 has van der Waals' forces, whilst HCl has permanent dipole–dipole attractions. Permanent dipole–dipole attractions are much stronger than induced dipoles. More energy is required to break the stronger permanent dipole–dipole attractions, and therefore HCl has a higher boiling point.

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