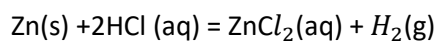


Answer on Question 82991 in General Chemistry

$$.m(\text{Zn}) = 16.0 \text{ g}$$



$$.m(\text{H}_2) = ?$$

Find the amount of substance of Zn

$$.n(\text{Zn}) = \frac{m}{Ar(\text{Zn})} = \frac{16}{65} = 0.246 \text{ mol}$$

$$.n(\text{H}_2) = n(\text{Zn}) = 0.246 \text{ mol}$$

$$.m(\text{H}_2) = n \times Mr = 2 \times 0.246 = 0.492 \text{ g}$$

$$Mr(\text{H}_2) = 2 Ar(\text{H}) = 2 \times 1 = 2$$

Answer provided by www.AssignmentExpert.com