

What is the mass in grams of  $5.0 \times 10^{20}$  molecules of aspirin ( $C_9H_8O_4$ ) ?

**Solution.**

$$N/N_A = m/M$$

$$m = NM/N_A = 9 \times 10^{22} / 6.02 \times 10^{23} = 0.15 \text{ g}$$

**Answer. 0.15 g**

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