

Answer on Question #82885 - Chemistry - Physical Chemistry

Question:

Which of the following statement about enthalpy is incorrect

- A. Enthalpy and internal energy of a system are always identical
- B. Enthalpy is a state of function
- C. ΔH is the enthalpy change at constant pressure
- D. ΔH is the enthalpy change at constant pressure

Reactions which absorb heat have a positive ΔH

Solution:

Entropy is a function of the state of the thermodynamic system and is widely used in thermodynamics.

So, incorrect statement is B. Enthalpy is a state of function.

Answer: Enthalpy is a state of function

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