The complete mechanism for a reaction is considered to occur in two steps, one of which is slow and the other fast.
$A+B \rightarrow C+D$ slow
$A+C \rightarrow E+F$ fast

The stoichiometric equation for the reaction is:
$2 \mathrm{~A}+\mathrm{B}=\mathrm{D}+\mathrm{E}+\mathrm{F}$
$A+B+C=D+E+F$
$A+B=C+D$
$A+B=D+E+F$
$A+C=E+F$
Solution:
The stoichiometric equation for the reaction is $\mathrm{A}+\mathrm{B}+\mathrm{C}=\mathrm{D}+\mathrm{E}+\mathrm{F}$

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