The complete mechanism for a reaction is considered to occur in two steps, one of which is slow and the other fast.

 $A + B \rightarrow C + D$ slow

 $A + C \rightarrow E + F$ fast

The stoichiometric equation for the reaction is:

2A + B = D + E + F

A + B + C = D + E + F

A + B = C + D

A + B = D + E + F

A + C = E + F

Solution:

The stoichiometric equation for the reaction is A + B + C = D + E + F

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