If you reacted 88.9 g of ammonia with excess oxygen, what mass of water would you expect to make? You will need to balance the equation first.
$NH_3(g) + O_2(g) \rightarrow NO(g) + H_2O(g)$

Solution:

$4NH_3(g) + 5O_2(g) \rightarrow 4NO(g) + 6H_2O(g)$
$M(NH_3) = 17$ g/mol

$$n = \frac{m}{M}$$

n = 5,23 mol

According to equation

$n(H_2O) = 6x5,23/4=7,845$ mol

$$m = n * M$$

$M(H_2O) = 18$ g/mol

$m(H_2O) =18$ x7,845=141,2 g

Answer: $m(H_2O) = 141,2$ g