

what kind of intermolecular forces are present in the following substances and explain in 1 sentence:

a. NH_3

b. SF_6

c. PCl_3

d. LiCl

e. HBr

f. CO_2

g. H_2S

h. CH_3OH

i. C_2H_6

Answer: in the compounds: NH_3 , SF_6 , PCl_3 , HBr , CO_2 , C_2H_6 , where there are covalent nonpolar and covalent polar bonds, as well as, for molecules of ammonia and sulfur hexafluoride there are unshared electron pairs, intermolecular bonds will be formed due to van der Waals interaction, that is, due to the coordinated motion of electrons along the orbitals. In the compound: LiCl , where an ionic bond, will be an intermolecular electrostatic interaction, between ions of different signs. In the compounds: H_2S , CH_3OH , where there is hydrogen and more electronegative elements: sulfur and oxygen, the intermolecular bond will have the character of hydrogen bonding, as well as van der Waals interaction.

Answer provided by AssignmentExpert.com