

#79856 Chemistry, General Chemistry

Explain in 2 to 3 sentences

1. What factors cause hydrogen bonds to be so much stronger than typical dipole-dipole bonds?
2. What kinds of intermolecular forces are present in the following substances:
 - a. H_2S
 - b. CH_3OH
 - c. C_2H_6

Answer:

1. Dispersion forces and large change in electronegativity cause hydrogen bonds to be so much stronger than typical dipole-dipole bonds.
2. a. H_2S – dipole-dipole
b. CH_3OH - London dispersion forces and hydrogen bonding
c. C_2H_6 - London dispersion forces.

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