

explain in 2 to 3 sentences.

1. Suggest some ways that the dipoles in London forces are different from the dipoles in dipole-dipole attractions.

2.A. Which would have a lower boiling point: O_2 or F_2 ? Explain

B. Which would have a lower boiling point: NO or O_2 ? Explain

3. Which would you expect to have a higher melting point (or boiling point): C_8H_{18} or C_4H_{10} ? Explain

Answer:

1. Dipoles in London forces are present in all compounds, can occur between atoms or molecules, are due to electron movement, are transient in nature (dipole-dipole are more permanent). London forces are weaker.
2. A. O_2 would have a lower boiling point, because F_2 have stronger London attractions due to molecule's size.
B. NO would have a lower boiling point, because O_2 has only London forces.
3. C_8H_{18} would have the higher melting/boiling point. C_4H_{10} is smaller, that's why it has fewer sites available for London forces to form.

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