Answer:

To estimate the lattice enthalpy of sodium chloride, we use $z(Na^+) = +1$, $z(Cl^-) = -1$, $A = 1.748$ and $d = r_{Na+} + r_{Cl-} = 283$ pm, hence (using fundamental constants from inside the back cover and ensuring that the units of d are appropriate to each part of the equation)

$$\Delta_L H^\ominus = \frac{(6.022\times10^{23}\,\text{mol}^{-1})\times|(+1)\times(-1)|\times(1.602\times10^{-19}\,\text{C})^2}{4\pi\times(8.854\times10^{-12}\,\text{J}^{-1}\text{C}^2\text{m}^{-1})\times(2.83\times10^{-10}\,\text{m})}\times\left(1-\frac{34.5\text{pm}}{283\text{pm}}\right)\times1.748$$

$$= 7.56\times10^5\,\text{J mol}^{-1}$$

or 756 kJ mol^{-1}.