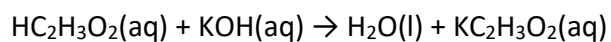


Answer on Question #77881, Chemistry / Organic Chemistry

Question:

What is the molarity of the acetic acid solution if 25.7 mL of a 0.240 M KOH solution is required to titrate 29.0 mL of a solution of HC₂H₃O₂?



Express your answer with the appropriate units.

Solution:

Amount of KOH: $0.240 \cdot 0.0257 = 0.006168 \text{ mol}$

Amount of HC₂H₃O₂: 0.006168 mol

Molarity of HC₂H₃O₂: $0.006168 / 0.0290 = 0.2127 \text{ mol/L} = 0.2127 \text{ M}$

Answer:

$0.2127 \text{ mol/L} = 0.2127 \text{ M}$

Answer provided by AssignmentExpert.com