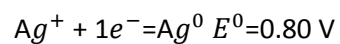


Answer on Question 77585 in General Chemistry



Find the cell voltage of the galvanic cell  $\text{Mn}|\text{Mn}^{2+}(\text{aq})||\text{Ag}^+(\text{aq})|\text{Ag}$

$$E = E^0(\text{Ag}^+/\text{Ag}^0) - E^0(\text{Mn}^{2+}/\text{Mn}^0) = 0.80 - (-1.18) = 1.98 \text{ V}$$

True statements are 1 and 2

1. The cell voltage will be 1.98 V
2. One half-cell reaction is  $\text{Mn} = \text{Mn}^{2+} + 2e^-$

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