## Question #77583, Chemistry / Physical Chemistry

According to a table of standard reduction potentials, what is the best reducing agent among the following species?

Pb

 ${\rm Mg}^{2+}$ 

 $Ag^{+}$ 

Ag

 $H_2$ 

## **Solution:**

According to a table of standard reduction potentials ( $\underline{\text{http://www.chemeddl.org/services/moodle/media/QBank/GenChem/Tables/EStandardTable.htm}$ ):

Pb

 $2H^{+}(aq) + 2e - = H_{2}(g) E = 0$ 

 $Pb^{2+}(aq) + 2e - = Pb (s) E = -0.126$ 

Mg<sup>2+</sup> cant be further oxidated

Ag<sup>+</sup> cant be further oxidated

 $Ag^{+}(aq) + e^{-} = Ag(s) E = +0.7994$ 

## Answer:

Pb

Answer provided by AssignmentExpert.com