

Answer on the question 77504 Chemistry / Inorganic Chemistry

$$0,453\text{V} = 0.799\text{V} + \frac{0.059}{1}\lg[\text{Ag}^+]$$

$$0.059\lg[\text{Ag}^+] = 0.453 - 0.799$$

$$\lg[\text{Ag}^+] = -5.86$$

$$[\text{Ag}^+] = 1.38 \cdot 10^{-6}$$

$$K_{\text{sp}} = [\text{Ag}^+]^2 \cdot [\text{C}_2\text{O}_4^{2-}] = (1.38 \cdot 10^{-6})^2 \cdot 0.33 = 6.3 \cdot 10^{-13}$$

Answer provided by AssignmentExpert.com