$0,453V = 0.799V + \frac{0.059}{1}lg[Ag^+]$

$0.059lg[Ag^+] = 0.453\text{-}0.799$

$lg[Ag^+] = \text{-}5.86$

$[Ag^+] = 1.38 \cdot 10^{-6}$

$K_{sp} = [Ag^+]^2 \cdot [C_2O_4^{2-}] = (1.38 \cdot 10^{-6})^2 \cdot 0.33 = 6.3 \cdot 10^{-13}$