

## Answer on Question #77111 – Chemistry – Other

### Task:

Benzene,  $C_6H_6$  is a potent carcinogen. How many grams are in 0.250 mole sample of benzene?

### Solution:

*Molar mass of benzene* =  $M(C_6H_6) = 6 * Ar(C) + 6 * Ar(H)$ ;

$$M(C_6H_6) = 6 * 12 + 6 * 1 = 78 \text{ g/mol}$$

Use the compound's molar mass:

$$0.25 \text{ mol} * \frac{78 \text{ g}}{\text{mol}} = 19.5 \text{ g of benzene}$$

**Answer:** 19.5 grams are in 0.250 mole sample of benzene