Answer on Question #75283, Chemistry / Inorganic Chemistry

1. Reaction showing acidic behaviour of sulphur dioxide:

$$SO_2 + KOH \rightarrow K_2SO_3 + H_2O$$

2. Reaction depicting the sulphonation by H<sub>2</sub>SO<sub>4</sub>

$$2H_2SO_4 \rightarrow SO_3 + H_3O^+ + HSO_4^-$$

3. Reactions of lead chambers during the manufacture of H2SO4 by Lead Chamber process

$$2ONOSO_2OH + H_2O \rightarrow 2H_2SO_4 + NO + NO_2$$
 
$$SO_2 + NO_2 + H_2O \rightarrow H_2SO_4 + NO$$

4. Reaction distinguishing between peroxomonosulphuric acid and peroxodisulphuric acid

$$5H_2SO_3(O_2) + 2C_6H_6 = 5SO_2 + 2CO_2 + 10C + 11H_2O$$
.

5. Reduction reaction of H<sub>2</sub>O<sub>2</sub> with chlorine

$$H_2O_2 + Cl_2 \rightarrow O_2 + 2HCl$$

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