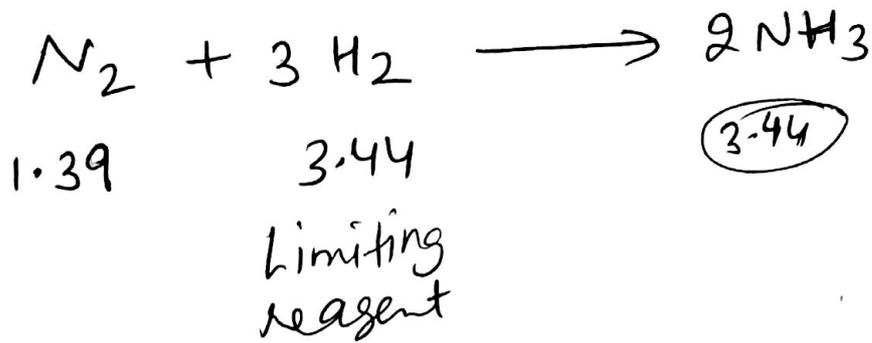


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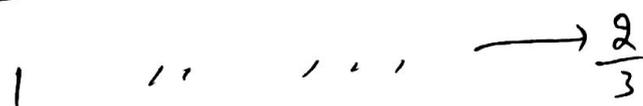
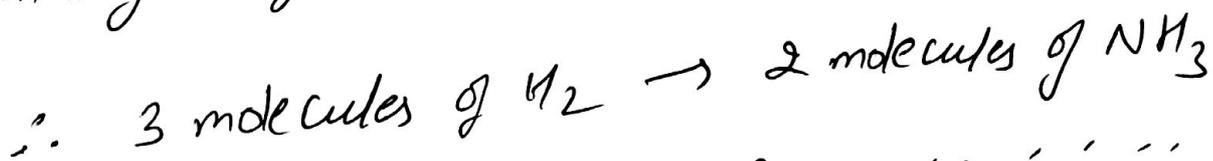


Acc. to stoichiometric eqⁿ



$$1.39 \longrightarrow 3 \times 1.39 = \underline{4.17 \text{ moles}}$$

So, limiting reagent is (H_2).



$$= \underline{2.293 \text{ moles}}$$

Now, moles in grams of NH_3 is:-

$$= 2.293 \times 17$$

$$= \boxed{38.981 \text{ g}}$$

[\because Molecular Mass of $\text{NH}_3 = 17$]