#74962

$$N_2 + 3H_2 \longrightarrow 2NH_3$$

1.39 3.44 (3.44)

Limiting reagent

ACC. to stoichiometric Eqn

$1 N_2 \longrightarrow 3 H_2$

$1.39 \longrightarrow 3 \times 1.39 = 4.17$ moles

So, Limiting reagent is (H_2).

$\therefore$ 3 molecules of H_2 $\longrightarrow$ 2 molecules of NH_3

1 " " $\longrightarrow \frac{2}{3}$ " "

3.44 " " $\longrightarrow \frac{2}{3} \times 3.44$

$= 2.293$ moles

Now, moles in grams of NH_3 is:-

$= 2.293 \times 17$ [$\because$ Molecular Mass of $NH_3 = 17$]

$= \boxed{38.981\ g}$

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