#74881

① $AgBr_{(s)} \rightleftharpoons Ag^+_{(aq)} + Br^-_{(aq)}$

The K_{sp} expression is :-

$K_{sp} = \{Ag^+\} \{Br^-\}$

$= 7.1 \times 10^{-7} \times 7.1 \times 10^{-7}$

$\boxed{K_{sp} = 50.41 \times 10^{-14}}$ ⇐

② $BaF_{2(s)} \rightleftharpoons Ba^{2+}_{(aq)} + 2F^-_{(aq)}$

The K_{sp} expression is :-

$K_{sp} = [Ba^{2+}] [F^-]^2$

$= 1.5 \times 10^{-6} \times 4 \times (1.5 \times 10^{-6})^2$

$= 13.5 \times 10^{-18}$

$\boxed{K_{sp} = 135 \times 10^{-19}}$ ⇐