

Answer on Question # 73649 - Chemistry - Other

Calculate the empirical formula of vulcium chloride.

Answer: there is no such element as vulcium (Vu). In order to get the numerical answer, the atomic weight of Vu and composition of vulcium chloride are needed. To calculate the empirical formula in general, the following steps are required:

1. Convert percent composition to mass, assuming 100 g of the substance (assume that there are 100 g of vulcium chloride and then find the mass of vulcium and chlorine in this sample).

2. Convert mass to moles, using the molar masses of the elements:

moles = grams/(molar mass).

3. Find the mole ratio of the elements present in a compound. For example, if the mole ratio of Vu:Cl is 3:2, then the empirical formula is Vu_3Cl_2 .

Sources:

<https://www.chemteam.info/Mole/EmpiricalFormula.html> (ChemTeam: Empirical Formula)

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