

## Answer on Question #71796 – Chemistry – Other

### Task:

If a solution of aspirin has a  $[\text{OH}^-] = 4.4 \times 10^{-12}$  M, what is the pH of the solution?

### Solution:

$$pH + pOH = 14;$$

$$pOH = -\log[\text{OH}^-] = -\log(4.4 \times 10^{-12}) = 11.3565 \approx 11.36;$$

$$pH = 14 - pOH = 14 - 11.36 = 2.64;$$

$$pH = 2.64$$

**Answer:** pH=2.64

Answer provided by <https://www.AssignmentExpert.com>