

### Answer on Question # 69246, Chemistry / General Chemistry

A 0.745 gram sample of potassium ferricyanide ( $K_3Fe(CN)_6$ ) requires 15.43 ml of thiosulfate solution to reach the iodometric end point. Calculate the Molarity of the thiosulfate solution.

**Solution:**

$$c(Na_2S_2O_3) = \frac{1000 * m(K_3 Fe (CN)_6)}{V(Na_2S_2O_3) * E(K_3 Fe (CN)_6)}$$

$$c(Na_2S_2O_3) = \frac{1000 * 0.745}{15.43 * 329.25} = 0.1466 M$$

**Answer:** 0.1466 M.

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