## Answer on Question \#69149 - Chemistry - Other

## Task:

Balance the following redox reaction. When this equation is properly balanced, using the smallest whole number coefficients, what are the coefficients on each substance?

$$
\mathrm{Al}(\mathrm{~s})+\mathrm{Cu}^{2+}(\mathrm{aq})=\mathrm{Al}^{3+}(\mathrm{ag})+\mathrm{Cu}(\mathrm{~s}) .
$$

## Solution:

The reaction equation:

$$
\mathrm{Al}+\mathrm{Cu}^{2+}=\mathrm{Al}^{3+}+\mathrm{Cu} ;
$$

$A l^{0}-3 e=A l^{3+} ;$
$\mathrm{Cu}^{2+}+2 e=\mathrm{Cu}^{0} \mid ;$
$2 \mathrm{Al}+3 \mathrm{Cu}^{2+}=2 \mathrm{Al}^{3+}+3 \mathrm{Cu} ;$
$2,3,2,3$
$\mathrm{Cu}^{2+}$ - oxidizing agent;
Al-reducing agent;
Answer: B. 2,3,2,3.
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