$Mg^{2+} + 2OH^- \leftrightarrow Mg(OH)_2$

The solution constant K_s :

$K_s = [Mg^{2+}] \times [2OH^-]^2 = s \times (2s)^2 = 4s^3 = 1.5 \times 10^{-12}$ ($mol^3 l^3$)

$s = \sqrt[3]{1.5/4} \times 10^{-4} = 0.7211 \times 10^{-4} = 7.21 \times 10^{-5}$ ($mol \times l$)

The concentration of $[Mg^{2+}] = 7.21 \times 10^{-5}$ $mol \times l$, and the concentration of $[OH^-] = (1.5 \times 10^{-12})/(7.21 \times 10^{-5}) = 2.08 \times 10^{-8}$ $mol \times l$.