

## Answer on Question # 60479 - Chemistry - Organic Chemistry

What rule is violated in the following electronic configuration  $1s^1 1s^2 2p_x^2 2p_y^1 2p_z^0$ .

### Solution

According to the Hund's rule, every orbital in a subshell should be occupied with one electron before any other orbital from this subshell is doubly occupied. The electronic configuration given for p subshell represents two electrons in  $p_x$  orbital, whereas the  $p_z$  orbital is empty. It is not in compliance with the Hund's rule. The correct electronic configuration should be

$1s^1 1s^2 2p_x^1 2p_y^1 2p_z^1$ .

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