1. Lead (ll) sulphate is used as white pigment. what is the solubility of $PbSO_4$? ksp is $1.8x10^{-11}$

Solution:

Ksp $PbSO_4$ = $[Pb^{2+}][SO_4^{2-}]$
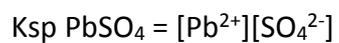
S – solubility

$[Pb^{2+}]$ = $[SO_4^{2-}]$ = S

Ksp = S×S=S^2

S = $\sqrt{Ksp}$

S = $\sqrt{1.8 \times 10^{-11}}$ = $4.24×10^{-6}$

Answer: S = $4.24×10^{-6}$

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