

## Answer on the question 59122, Chemistry, Other

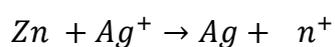
### Question:

what is the ionic equation of silver nitrate solution and zinc metal

### Answer:

Zinc metal has less reduction potential (-0.76 V) than silver (+0.80 V). This means, that silver is more stable in its reduced (metallic form) than zinc.

Silver nitrate solution contains silver and nitrate ions:  $\text{Ag}^+$  and  $\text{NO}_3^-$ . The reaction with zinc gives zinc cation and reduced silver as products:



The full ionic equation of this reaction is:

