

**Question #58738, Chemistry / Inorganic Chemistry**

A solution is prepared by dissolving 4.40 g of KSCN in enough water to make 250 mL of solution. What is the concentration of the solution?

**Answer**

$$C_M = \frac{v}{V};$$

$$v = \frac{m}{Mr} = \frac{4.4}{97} = 0.045mole$$

$$C_M = \frac{0.045mole}{0.25L} = 0.18mole/L$$

The concentration of the solution is 0.18 mole/L or 0.18 M