

## Answer on the Question #57934

### What makes hydrogen peroxide a polar molecule

Basically, we can call a molecule polar, if its electrons are unequally distributed. In the hydrogen peroxide, H and O are bond by covalent polar bond, which means that electron pairs are localized closer to the oxygen. As a result, oxygen atoms carry  $-\delta$  charge, and hydrogen  $+\delta$  charge. Due to the 3D structure (it is bent and rotated), there is a positive pole and a negative pole in the molecule. And that's why  $\text{H}_2\text{O}_2$  is polar.

