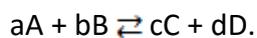


Answer on Question#57575 – Chemistry – General chemistry

Question: What are K_p and K_c? What is the relation between them?

Answer: K_p and K_c are equilibrium constants.

Example, for reaction:



$$K_c = \frac{[C]^c \cdot [D]^d}{[A]^a \cdot [B]^b}$$

Where [] is equilibrium concentration.

$$K_p = \frac{P_C^c \cdot P_D^d}{P_A^a \cdot P_B^b}$$

Where P is partial pressure

Relation between this constants:

$$K_p = K_c(RT)^{\Delta n}$$

Where R is universal gas constant; Δn is (number of moles of gaseous products) – (number of moles of gaseous reactants).