

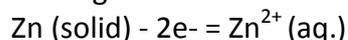
Answer on Question #56949 - Chemistry - Other

Question:

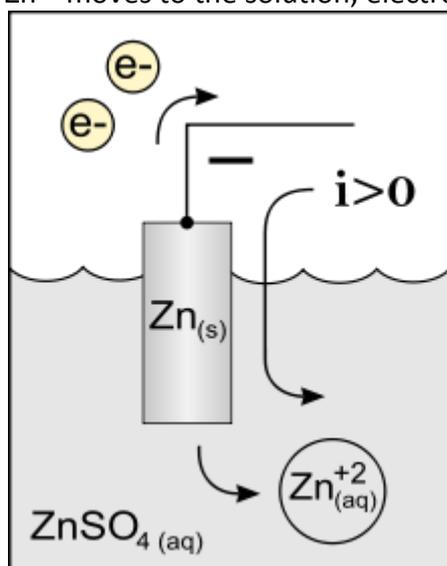
Describe the direction of ion and electron flow in a anode/cathode diagram (galvanic cell diagram)

Answer

Anode is negative electrode in an electrolytic cell where oxidation occurs. For example we have Zn anode. Zn as an active metal dissolves and gives electrons:

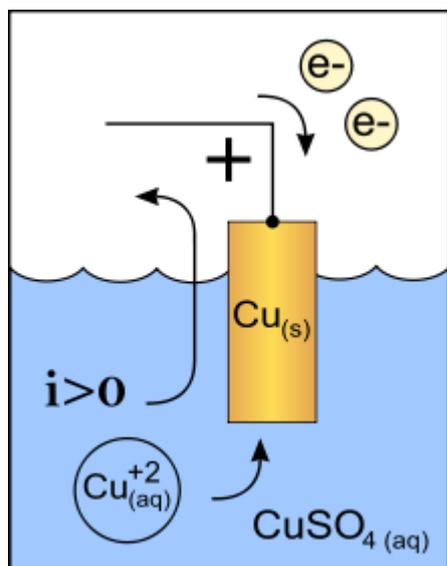
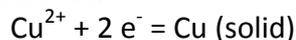


Zn^{2+} moves to the solution, electrons – from anode to cathode of galvanic cell.



For example, we have copper cathode. Cathode is a positive electrode in electrolytic cell.

Copper ions from the water solution of it's salt (CuSO_4) moves to cathode and it forms copper-metal with help of that electrons from anode:



All the pictures from WIKI: https://en.wikipedia.org/wiki/Galvanic_cell.