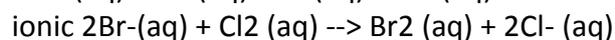
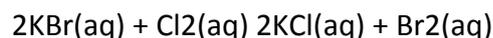
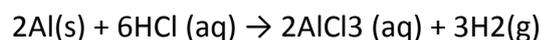
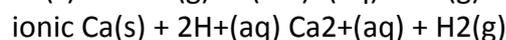
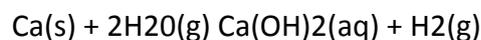
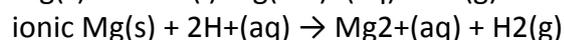
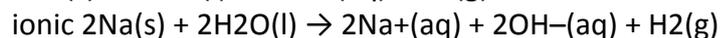
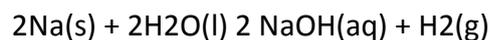
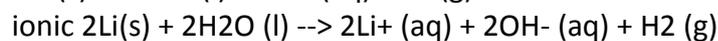
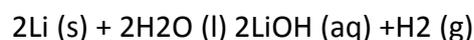


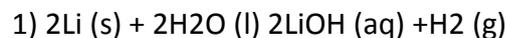
## Answer on Question #56294 - Chemistry - General Chemistry

### Question:

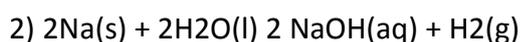
would these equations and ionic equations be correct along with the states thank you



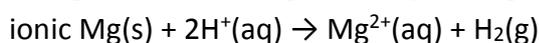
### Answer:



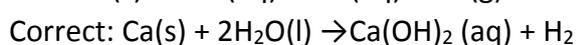
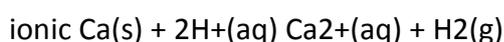
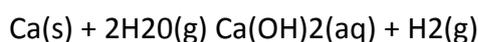
Correct!



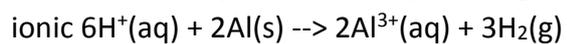
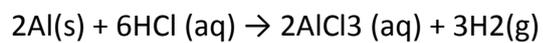
Correct!



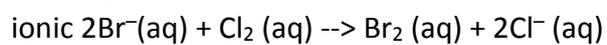
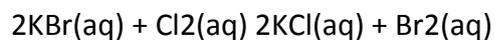
But,  $\text{Mg(OH)}_2$  is only sparingly soluble in water, forming the suspension. Anyway, some quantity of ions is also formed. Thus, it will be correct to write also  $\text{Mg(OH)}_2\text{(aq)}$ , and  $\text{Mg(OH)}_2 \text{ (s)}$ , dependently on the condition.



The situation with Calcium hydroxide is similar with the situation of Magnesium. The difference is lower solubility of  $\text{Ca(OH)}_2$  in water. Thus, more correct is to write the product in the form of  $\text{Ca(OH)}_2\text{(aq)}$ .



Correct! Talking about the interaction of Aluminum with water, one can note thick Aluminum oxide layer that appears when the Aluminum is exposed to water. This layer is slowing down the reaction. Once the layer is corroded, the hydrogen gas forms, according to the reaction above.



Correct!