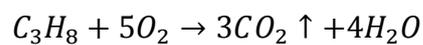


Answer on Question #54891 - Chemistry - General chemistry

Question:

How many grams of $O_2(g)$ are needed to completely burn 87.3g of $C_3H_8(g)$?



$$\frac{m_{C_3H_8}}{M_{C_3H_8}} = \frac{m_{O_2}}{5 \times M_{O_2}}$$

$$\frac{87.3}{44.1} = \frac{m_{O_2}}{5 \times 32}$$

$$m_{O_2} = \frac{87.3 \times 5 \times 32}{44.1}$$

$$m_{O_2} = 316.73g$$

Answer: 316.73g