

Answer to Question #53906 – Chemistry – General chemistry

Question:

If the heat of vaporization of water at 25 degree Celsius is 44.0 kJ/mole how much heat does your body lose when you sweat off 5.00 g of water at room temperature?

Solution:

$$Q = \frac{m \times \Delta H}{M_r} = \frac{5.00 \text{ g} \times 44.0 \text{ kJ/mole}}{18.00 \text{ g/mole}} = 12.22 \text{ kJ}$$

Answer:

12.22 kJ