

## Answer on Question #53621 – Chemistry – General Chemistry

### Question:

How many moles of NaOH are present in 13.0 mL of 0.190 M NaOH?

### Answer:

Molar concentration or molarity is most commonly expressed in units of moles (solute) per litre (solution):

$$c = \frac{n}{V}$$

Here, **c** is the amount of the solute in moles, **n** is the number of moles present in the volume **V** (in litres).

That's why we can write the following expression:

$$n = c \cdot V = 0.190 \cdot 0.013 = 0.00247 \text{ moles}$$

**Answer: 0.00247 mol.**