Let’s first convert grams of water to moles of water (molar mass of water is 18 g/mol):

$$1255g\ H_2O \cdot \left(\frac{1mol}{18.0g}\right) = 69.7mol\ H_2O.$$

The amount of heat released when $1mol$ of vapor condenses is called a molar heat of condensation ($\Delta Hcond$):

$$\Delta Hvap = -\Delta Hcond.$$

The molar heat of vaporization of water is $\Delta Hvap = 40.7\ kJ/mol$.

Then, $\Delta Hcond = -40.7\ kJ/mol$ and the next step is a convertion from moles of water to ΔH, multiplying by the $\Delta Hcond$:

$$\Delta H = 69.7mol\ H_2O \cdot \frac{-40.7kJ}{1mol\ H_2O} = -2837kJ.$$

The negative sign indicate that heat is given off.

Therefore, the process will release $\boldsymbol{2837kJ}$ of heat.