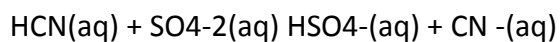


Answer on the question #50447, Chemistry, Other

Question:

Why hydrogen ions are NEVER found in an aqueous solution?



What is the Bronsted - Lowry acid in this equation?

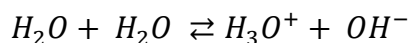
What is the Bronsted - Lowry base in this equation?

What is the conjugate acid in this equation?

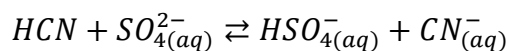
What is the conjugate base in this equation?

Answer:

- 1) The hydrogen ions are never found in the solutions because they always exist in the H_3O^+ form. This originates from the true water autoionisation reaction:



2)



the Bronsted - Lowry acid is the HCN, because it donates hydrogen ion;

the Bronsted - Lowry base is $\text{SO}_4^{2-}(\text{aq})$, because it accepts the hydrogen ion;

the conjugate acid is the $\text{HSO}_4^-(\text{aq})$, the conjugate base – $\text{CN}^-(\text{aq})$.