

Answer on Question #50375, Chemistry, Other

What is the pOH of a solution with a [OH⁻] concentration of 9.7 X 10⁻¹¹ M?

Solution:

$$pOH = -\log_{10}[OH^{-}]$$

$$pOH = -\log_{10} 9.7 \times 10^{-11} = 10.01$$

Answer:

pOH = 10.01

<https://www.AssignmentExpert.com>