

Answer on Question #50374, Chemistry, Other

Find the pH of a solution whose pOH is 5.36.

**Solution:**

$$\begin{aligned}[H^+] \times [OH^-] &= 10^{-14} \\ \log_{10}[H^+] + \log_{10}[OH^-] &= -14 \\ pH + pOH &= 14 \\ pH &= 14 - pOH \\ pH &= 14 - 5.36 = 8.64\end{aligned}$$

**Answer:**

**pH is 8.64**

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